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| Subject | Chemistry |
| Unit/Topic | Year 10 Quantitative Chemistry |

| Key Vocabulary | Definition |
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| Avogadro Constant | The number of particles in one mole of particles, 6.02×10^{23} per mole. |
| Balanced Chemical Equation | A chemical equation written using the symbols and formulae of the reactants and products, so that the number of units of each element present is the same on both sides of the arrow. |
| Concentration | The concentration of a solution tells us how much of a substance is dissolved in water. The higher the concentration, the more particles of the substance are present. |
| Dissolve | When a substance breaks up and mixes completely with a solvent to produce a solution. |
| Excess | In chemistry, a substance is in excess if there is more than enough of it to react with another reactant. |
| Formula | A combination of symbols that indicates the chemical composition of a substance. |
| Limiting Reactant | The reacting substance that is completely used up in a chemical reaction and which determines how much product is made. |
| Mass | The amount of matter an object contains. Mass is measured in kilograms (kg) or grams (g). |
| Mole | The amount of substance that contains the same number of particles as there are atoms in 12 g of carbon-12 (contains the Avogadro's constant 6.0×10^{23} number of particles). |
| Molecule | A collection of two or more atoms held together by chemical bonds. |
| Product | A substance formed in a chemical reaction. |
| Reactant | A substance that reacts together with another substance to form products during a chemical reaction. |

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| Relative Atomic Mass | The mean relative mass of the atoms of the different isotopes in an element. It is the number of times heavier an atom is than one-twelfth of a carbon-12 atom. |
| Relative Formula Mass | The sum of the relative atomic masses of the atoms in a chemical formula. |
| Resolution | For a measuring instrument, the smallest change in a quantity that gives a change in the reading that can be seen. |
| Solute | The dissolved substance in a solution. |
| Solution | Mixture formed by a solute and a solvent. |
| Solvent | The liquid in which the solute dissolves to form a solution. |
| Thermal Decomposition | Type of reaction in which a compound breaks down to form two or more substances when it is heated. |
| Uncertainty | The interval within which the true value of a quantity can be expected to lie. |
| Volume | The volume of a three-dimensional shape is a measure of the amount of space or capacity it occupies, e.g. an average can of fizzy drink has a volume of 330 ml. |