Subject	**Chemistry**
Unit/Topic	Year 10 Bonding and Structure

Key Vocabulary	Definition
Allotrope	Different structural forms of the same element in the same state. For example carbon has four allotropes – diamond, graphite, graphene and fullerenes.
Alloys	An alloy is a mixture of two or more elements, at least one of which is a metal.
Aqueous	Dissolved in water to form a solution. Shown as (aq) in chemical equations.
Atom	The smallest part of an element that can exist.
Brittle	If something is brittle it is easily broken.
Dense	The number of particles per volume. If you have a high density your substance will be heavier than one with a lower density.
Ductile	The ability of a substance to be stretched into wires.
Gas	State of matter with weak forces of attraction between particles. Particles move randomly. Gases have no fixed volume or shape.
Intermolecular	Weak attractive forces between molecules. When a simple molecular substance melts or boils, it is the intermolecular forces that are broken (not the covalent bonds in each molecule).
Intramolecular	Forces within molecules. For example strong covalent bonds within water molecules.
Ion	Electrically charged particle, formed when an atom or molecule gains or loses electrons.
Liquid	State of matter where randomly arranged particles tend to stickly close together but can move past one another. Liquids have a fixed volume but no fixed shape.
Malleable	Capable of being hammered or pressed into a new shape without being likely to break or return to the original shape.

Polymer	A large molecule formed from many identical smaller molecules known as monomers.
Solid	State of matter where particles are held close together with strong forces of attraction to form a regular lattice arrangement. Solids have a fixed shape and volume.
Sonorous	Able to produce a deep or ringing sound. For example metals.